



Chemistry

Name: _____

Section _____ EMPIRICAL FORMULAS Date: _____

Determining empirical and molecular formulas.

1. The empirical formula of a compound is NO_2 . Its molecular mass is 92 g/mol. What is its molecular formula?

2. The empirical formula of a compound is CH_2 . Its molecular mass is 70. g/mol. What is its molecular formula?

3. The empirical formula of a compound is CH_2O . Its gram formula mass is 60. g/mol. What is its molecular formula?

4. The empirical formula of a compound is $\text{C}_4\text{H}_{10}\text{O}$. Its gram formula mass is 74 g/mol. What is its molecular formula?

5. A compound with a molecular mass of 88 g/mol has an empirical formula of $\text{C}_2\text{H}_4\text{O}$. What is its molecular formula?

6. Write the empirical formula for C_6H_8 .

7. Write the empirical formula for $\text{C}_6\text{H}_{12}\text{O}_6$.
